

Name _____

Date _____

The Periodic Table of Elements Worksheet

Directions: Fill out this worksheet while watching the Power Point Presentation.

I. History

1. What did chemists recognized in 1800s about elements?
2. How did chemists try to organize elements?
3. What was unique about Mendeleev's table of elements?
4. Why did Mendeleev left extra spaces or holes in his table?
5. Why did Mendeleev's table become a great tool for scientists?
6. Mendeleev is credited as the _____ of modern _____ table.

II. Information

7. _____ is one or two letters that are used to represent the element's name.

8. Atomic number is _____

9. Elements are arranged in the order of _____ atomic number.

10. Atomic mass is the _____

11. The atomic mass _____ as the atomic number _____, but not evenly.

12. _____ is the smallest particle into which an element can be divided and still have the properties of that element.

13. _____ is surrounded by negatively charged electrons

14. The nucleus consists of particles called _____ and _____.

15. Protons are

16. Neutrons are

17. _____ are electrons that have the highest energy level and are held most loosely.

III. Organization of the Periodic Table

18. _____ is row of elements on the periodic table.

19. _____ or _____ is a column on the periodic c table.

IV. Trends of The Periodic Table

20. _____ is the amount of energy it takes to strip away the first valence electron.

21. _____ is the energy required to remove an electron from a single charge anion

22. _____ is a measure of how tightly an atom holds onto its valence electrons

23. _____ of an atom is the distance between nucleus of an atom and the outer most electron shell.

24. _____ of an atom is the tendency of an atom to lose an electron instead of gaining electrons in order to fill it's valance shells.